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# The Effects of an Applied Voltage on the Corrosion Characteristics of Dense MgO

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## Abstract

In a recent investigation Mills and Riaz [1] showed that industrial oxide refractory corrosion by liquid oxides could be changed by the application of a small voltage across the liquid oxide-refractory interface. They explained their result in terms of penetration of the refractory pores with liquid oxide. Their analysis of the corrosion effect was to some degree limited by the use of an industrial refractory material in the study. Further, it was not clear whether their findings were limited to solely industrial refractories or had wider ranging application to more dense ceramic type solid oxide systems. In this study, a simpler and more easily characterized solid oxide (dense MgO) has been used to examine the effects of an applied voltage on the solid oxide in a liquid oxide melt. The dissolution rate of an MgO ceramic in a CaO-SiO<sub>2</sub>-Al<sub>2</sub>O<sub>3</sub> and CaO-SiO<sub>2</sub>-Fe<sub>2</sub>O<sub>3</sub>-FeO-MgO liquid oxide composition at various applied voltages has been measured at 1540°C. It was found that the MgO corrosion in the CaO-SiO<sub>2</sub>-Al<sub>2</sub>O<sub>3</sub> system was insensitive to an applied voltage over the voltage range -0.5 to 0.3 V. In the CaO-SiO<sub>2</sub>-Fe<sub>2</sub>O<sub>3</sub>-FeO-MgO liquid oxide system the MgO corrosion rate showed a maximum at -0.45 V. This effect has been explained by considering the consequences of an applied voltage on the rate of Marangoni flow at the liquid oxide-refractory-gas interface and in turn, the flow effect on the rate of the mass transfer controlled MgO dissolution reaction.

**Keywords:** Slags, Liquid Oxide, Refractory, MgO, Electrode Potential

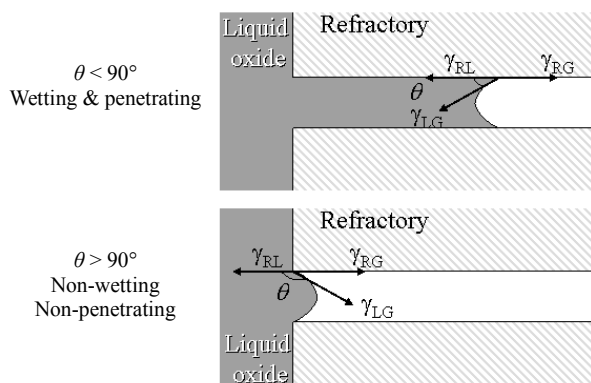
## 1. Introduction

MgO based refractories are widely used in the metallurgical industries and their service lifetimes are often limited by the rate the refractory dissolves in the liquid oxides generated in the processes [2]. The dissolution rate is a function of contact area between the liquid oxide and refractory which in turn is a function of the area calculated using the macro dimensions of the refractory and the pore structure within the refractory. Recently there have been two separate studies that have shown that the rate of liquid oxide penetration into a refractory and hence the rate of dissolution of a refractory into the liquid may be affected by the application of a small applied voltage across a liquid oxide-refractory interface [1,3]. The results of these studies were explained using an integrated form of Poiseuille's law [3], shown in Equation (1), for liquid penetration,  $l$ , into pores,

$$l^2 = \left( \frac{r\gamma \cos\left(\frac{\theta}{2}\right)}{\eta} \right) t \quad (1)$$

where  $r$  is the pore radius,  $\gamma$  is the interfacial tension,  $\theta$  the contact angle between the liquid oxide and refractory,  $\eta$  is the dynamic viscosity of the liquid oxide and  $t$  is time. In these studies it was argued that by application of an electrode potential across the refractory-liquid oxide interface the  $\gamma$  and  $\theta$  change thereby reducing the depth of penetration for a given time. A schematic showing liquid penetration in a wetting ( $\theta < 90^\circ$ ) and non-wetting ( $\theta > 90^\circ$ ) system are given in **Figure 1**.

The findings presented in these studies [1,3] require further investigation. Why should the interfacial tension and contact angle of a liquid oxide-refractory system change with changing applied voltage? Does this change in interfacial tension and contact angle of the liquid oxide-refractory system change the rates of dissolution of

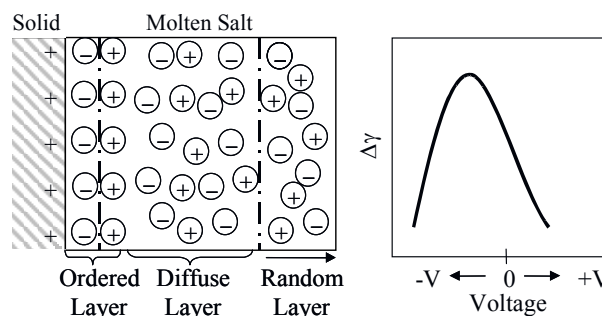


**Figure 1.** A schematic showing the effects of liquid penetration in a wetting and non-wetting system, where L = liquid oxide, R = refractory and G = Gas.

refractories with extremely low porosity where increased contact area generated by liquid oxide penetration into pores is not an issue? If their findings are to be fully understood with respect to refractory and ceramic corrosion then these question need to be addressed.

An answer to the first question may be found by considering the application of a number of existing theories that have attempted to explain the composition stratification in other systems such as metal-aqueous [4,5] solutions, metal-polymers [6] and molten salts-metal [5,7,8] systems under an applied voltage and the effects of this change on the interfacial tension of the system. In these systems it has been found that ions/molecules at or near the phase interface order themselves to minimize the voltage affects. This has the effect of changing the interfacial properties of the phase interface. A schematic of this ordering process is shown in **Figure 2**, for a molten salt system under an applied voltage. This ordering has the effect of changing the interfacial properties of the phase interface, shown as  $\Delta\gamma$  in **Figure 2**.

Highly basic liquid oxides are ionic in nature [9] and to all intents and purposes can be considered as a molten salt. Acidic liquid oxides are more covalent [9] and may not be easily represented as molten salts. It is likely that models, which describe the behavior of molten salts, will have some applicability to liquid oxides in general and basic liquid oxides in particular. Unfortunately, Kazakov *et al.* [3] and Riaz *et al.* [1] investigations were on penetration of liquid oxide into pores. It is difficult to define the contact area between a porous refractory and the liquid oxide in these experiments. This precludes direct application of molten salt theories and hence an answer to the question why should the interfacial tension and contact angle of a liquid oxide-refractory system change with changing applied voltage? The study of a simpler refractory-liquid oxide system under an applied voltage with a more dense refractory, (minimal porosity) and therefore a better defined contact area may enable an answer to the question. It may be expected, though it has



**Figure 2.** A schematic of the ordering in a molten salt under an applied voltage and the resulting change in interfacial tension  $\Delta\gamma$ . Circles with + and – denote an ion with the respective polarity.

to be confirmed, that changes in interfacial tension and/or wetting of the liquid oxide-refractory system due to an applied voltage would also affect the dissolution rate of more dense refractories. It is known that changing the interfacial tension and/or wetting will also change the interfacial (Marangoni) flow conditions [10] at the liquid oxide-refractory-gas interface. Marangoni flows are interfacial tension driven flows and are known to affect corrosion at the liquid oxide-refractory-gas interface [2,10]. Such a change in local flow conditions would have a significant effect on refractory-liquid oxide dissolution systems that are at least in part controlled by mass transport in the liquid oxide phase. The confirmation that an applied voltage affects the dissolution behavior of the simpler dense refractory-liquid oxide system is a necessary pre-cursor to addressing more fundamental issues of why the interfacial tension and/or wetting change with applied voltage. In this study the dissolution of a dense MgO refractory in contact with a CaO-SiO<sub>2</sub>-Fe<sub>2</sub>O<sub>3</sub>-FeO-MgO liquid oxide under an applied voltage has been studied and reported.

MgO was chosen as the refractory to be studied as it is a widely used refractory in the liquid metal processing industry and it is known that the dissolution of this refractory in a CaO-SiO<sub>2</sub>-Fe<sub>2</sub>O<sub>3</sub>-FeO-MgO system is at least in part controlled by mass transfer in the liquid oxide phase [11,12] and therefore will be significantly affected by Marangoni flows.

## 2. Experimental

The effect of an applied voltage across a solid dense MgO-liquid oxide interface on the dissolution of rate of MgO was studied. The composition of liquid oxides used are given in **Table 1**. A cylindrical piece of high density MgO was dipped in liquid oxide for 90 or 120 minutes, removed and water quenched. The resulting corroded MgO sample was then sectioned, ground and polished to a 1 μm finish for microscopy using standard ceramographic

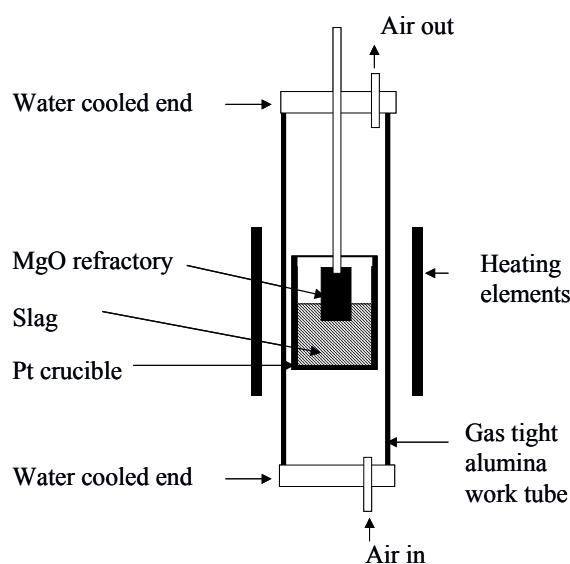
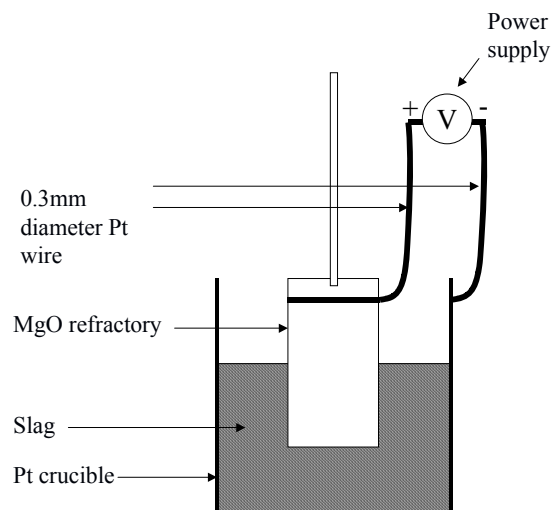
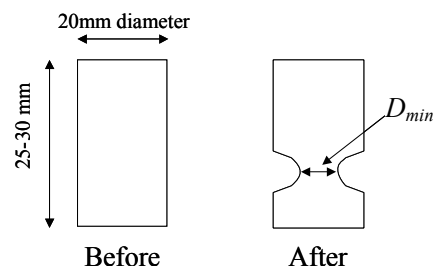
**Table 1. Liquid oxide compositions and dense MgO/liquid oxide contact times used in this study.**

Mass %	CaO	SiO <sub>2</sub>	Al <sub>2</sub> O <sub>3</sub>	MgO	Fe <sub>2</sub> O <sub>3</sub>	FeO	Time (min)
Liquid oxide 1	45.8	45.8	\	3.0	4.6	0.9	90
Liquid oxide 2	49.5	40.5	10	\	\	\	120

methods. The experiments were carried out in a dried air atmosphere and heated to 1540°C in a molybdenum disilicide vertical tube furnace. A schematic of the furnace set-up is given in **Figure 3**. A schematic of the electrical circuit used to apply the voltage is given in **Figure 4**. It should be noted that the electrical cell in **Figure 4** is analogous to a capacitance cell and that there is no current flowing.

The MgO dissolution was measured by measuring the minimum diameter,  $D_{min}$ , of the sample as defined in **Figure 5**. This change in diameter is a result of the refractory dissolution process, therefore the smaller  $D_{min}$ , the greater the amount of refractory dissolution/corrosion. The cylindrical MgO test samples were supplied by Rojan Advanced Ceramics, Henderson, Western Australia. They are greater than 97% pure and had a density of in excess of 96% theoretical.

The liquid oxide was prepared by pre-melting appropriate mixtures of CaO, SiO<sub>2</sub>, Fe<sub>2</sub>O<sub>3</sub> and MgO in air, quenching the liquid oxide and then crushing the resultant glass. This process was repeated to obtain a homogenous liquid oxide. The crushed oxide was then melted in-situ in the platinum crucible prior to running an experiment. The reported oxide compositions in **Table 1** are based on ICP measurements for CaO, SiO<sub>2</sub>,

**Figure 3. A schematic of the furnace set-up.****Figure 4. A schematic of the electrical circuit used to apply a voltage.****Figure 5. A schematic showing the MgO sample before and after the dissolution experiment showing liquid oxide line corrosion attack.  $D_{min}$  is the minimum diameter of the sample after corrosion.**

MgO, Al<sub>2</sub>O<sub>3</sub> and total Fe. A back titration method was used to establish the FeO content. The rest of the iron was assumed to be Fe<sub>2</sub>O<sub>3</sub>.

The solid MgO was chosen to be studied as it is a widely used refractory in the liquid metal processing industry and it is known that the dissolution of this refractory in a CaO-SiO<sub>2</sub>-Fe<sub>2</sub>O<sub>3</sub>-FeO-MgO system (Liquid oxide 1, **Table 1**) is at least in part controlled by mass transfer in the liquid oxide phase [11,12] and therefore will be significantly affected by Marangoni flows. The CaO-SiO<sub>2</sub>-Al<sub>2</sub>O<sub>3</sub> liquid oxide (Liquid oxide 2, **Table 1**) was chosen as there are published electrical capacitance data [13] available for this system. These data would aid future studies that assess models describing the ion behavior in liquid oxide systems under an applied voltage.

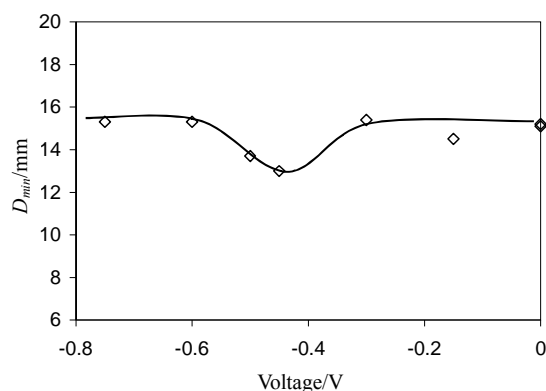
It was planned that the experimental immersion times for all experiments in this study would be constant. Unfortunately after 90 minutes using liquid oxide 2 there was minimal corrosion of the dense MgO. It was therefore decided to increase the experimental time for these experiments.

### 3. Results/Discussion

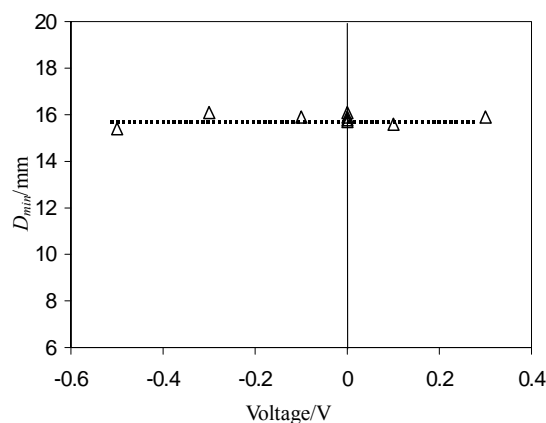
Experimental results for the effects of an applied voltage on the dissolution/corrosion of dense MgO in liquid oxides 1 and 2 are given in **Figures 6** and **7** respectively.

From inspection of **Figures 6** and **7** it can be seen that an applied voltage has a significant effect on the measured  $D_{min}$  when using liquid oxide 1 but not liquid oxide 2. At  $-0.5$  V using liquid oxide 2 there is a slight lowering of  $D_{min}$ . It is not clear whether this lowering of  $D_{min}$  using liquid oxide 2 is significant given the uncertainty in the data. The majority of the corrosion observed in this study takes place at the MgO refractory-liquid oxide-gas interface. At approximately  $-0.45$  V the amount of MgO corrosion using liquid oxide 1 and hence dissolution rate is at a maximum.

Why should an applied voltage across the refractory-liquid oxide interface affect the dissolution rate of the MgO refractory? It is known that MgO dissolution in the liquid oxide used in this study is at least in part controlled by mass transfer in the liquid oxide phase [11,12].

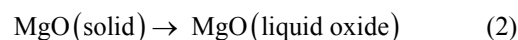


**Figure 6.** The effects of an applied voltage on the dissolution of dense MgO as characterized by  $D_{min}$  for liquid oxide 1 at  $1540^{\circ}\text{C}$  for 90 minutes.



**Figure 7.** The effects of an applied voltage on the dissolution of dense MgO as characterized by  $D_{min}$  for liquid oxide 2 at  $1540^{\circ}\text{C}$  for 120 minutes.

To discuss the effects of an applied voltage on a mass transfer controlled process it is instructive to consider a simple mass transfer model for steady state refractory corrosion, as shown in **Figure 8**. In this figure  $A$  and  $B$  represent the MgO in the refractory and liquid oxide phase respectively,  $C$  is the molar concentration and  $i$  represents interface. On the assumption that the dissolution process, as describe by Equation (2), is controlled purely by mass transfer in the liquid oxide



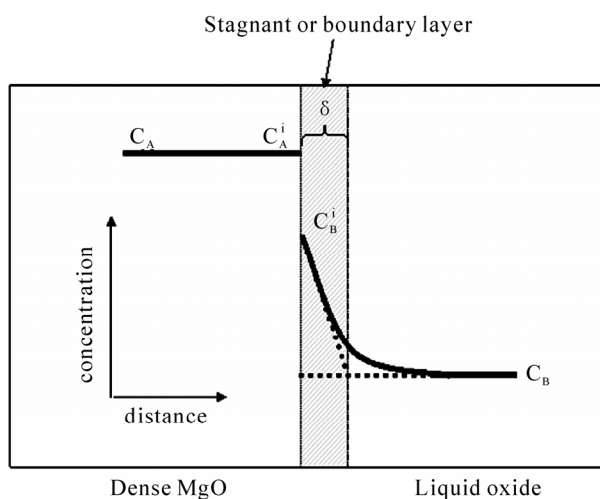
then the MgO dissolution flux,  $J$ , into the liquid oxide can be obtained from an application of Fick's law (Equation (3)) to the refractory-liquid oxide system

$$J = -\frac{D}{\delta} (C_{MgO} - C_{MgO}^i) \text{ mole.cm}^{-2}.\text{s}^{-1} \quad (3)$$

where  $D$  is the interdiffusion coefficient,  $\delta$  is the stagnant (boundary) layer as defined in **Figure 8**,  $C_{MgO}$  is the MgO concentration in the bulk of the liquid oxide and  $C_{MgO}^i$  is the MgO concentration in the liquid oxide at reaction interface.

Factors that effect the velocity of the liquid oxide adjacent the stagnant layer  $\delta$  will effect the concentration profile across the stagnant layer and hence the value of  $\delta$ . Any changes in  $\delta$  will change the MgO dissolution rate (flux  $J$  in Equation (3)). At the refractory-liquid oxide-gas interface it is known that Marangoni flows are active and play a significant part in the wear/corrosion profiles obtained in the refractory at that interface [14]. These flows are a result of shear stresses,  $\delta$ , caused by interfacial tension gradients and can be described by Equation (4) [14]

$$\tau = \frac{d\gamma}{dx} = \frac{\partial\gamma}{\partial T} \frac{dT}{dx} + \frac{\partial\gamma}{\partial C} \frac{dC}{dx} + \frac{\partial\gamma}{\partial\psi} \frac{d\psi}{dx} \quad (4)$$



**Figure 8.** A schematic of solid MgO dissolving in a liquid oxide showing concentration profiles on either side of the reaction interface.

where  $x$  is distance,  $T$  is temperature,  $C$  is concentration and  $\psi$  voltage. Other symbols are as defined previously. For all other things being equal, a change in the voltage term in Equation (4), as a result the applied voltage to the liquid oxide-refractory system will be reflected by a change in the Marangoni flow velocity. This in turn will affect the MgO dissolution flux, as the thickness of the stagnant layer  $\delta$  in Equation (3) will change. The MgO dissolution data as a function of applied voltage obtained in this study are consistent with, and therefore could be explained by such an argument. The argument that the applied voltage in the dense refractory-liquid oxide system studied in this investigation affects the interfacial phenomena, and hence the dissolution rate is consistent with Kazacov *et al.* [3] and Riaz *et al.* [1] explanation of the liquid oxide penetration observations of their porous refractory under an applied voltage. Essentially they argued that applying a voltage to their porous refractory-liquid oxide system changed the interfacial characteristics of the liquid oxide penetration.

The argument for the voltage effect on MgO for dissolution using liquid oxide 1 is based on the assumption that the MgO dissolution process is at least in part mass transfer controlled. Therefore providing MgO dissolution in a liquid oxide 2 is also at least in part mass transfer controlled a similar voltage effect should be observed. The results shown in **Figure 7** do not bear this out. Why there is this discrepancy is not clear and requires further study. It may be that the expected voltage effect takes place outside the range tested in this study, or that the particular liquid oxide composition used and ions formed are not as susceptible to and applied voltage effect, or that the voltage effect is not as significant in liquid oxide 2 as liquid oxide 1 and therefore masked by the uncertainty in the results, or that the dissolution process of MgO in liquid oxide 2 is not controlled by mass transfer in the liquid oxide phase.

It is unlikely that the dissolution process is not at least in part mass transfer controlled in the liquid oxide phase given previous work in this area [2,11,12], and more specifically [12] on the same solid liquid system, showed mass transfer control.

From the results of using liquid oxide 1 in this study and the work of Kazacov *et al.* [3] and Riaz *et al.* [1] it can be seen that the corrosion of a refractory may be altered by applying a voltage across the refractory-liquid oxide interface. This phenomenon offers the potential to be able to control, among other things, refractory corrosion in liquid metal production, a significant cost to many processes.

At present, it is not possible to predict with any certainty whether the applied voltage to the liquid oxide-refractory system will lessen or increase the dissolution rate of the refractory. Nor is it known with any certainty if any of the current models that have been devel-

oped to describe the behavior of a solid-liquid interface when a voltage is applied across the interface [4-8] are applicable to liquid oxide-refractory systems. These models describe how the ions stratify at the solid-liquid interface under an applied voltage and that the stratification is a function of temperature, liquid oxide composition (principally ion size) and liquid oxide capacitance.

Our understanding of ion sizes in steelmaking liquid oxides is primarily limited to basic (depolymerised) liquid oxides [9]. The ions present in acid liquid oxides, for example liquid oxides rich in silica, are not well described [9]. The lack of knowledge of acid liquid oxides is not necessarily a major problem as most liquid oxides used in liquid steel processing are basic.

Obtaining and utilizing liquid oxide capacitance data are key issues that must be addressed if a solution to the problems of direct application of the solid-liquid interface models to refractory dissolution that account for applied voltage effects are to be sought. There are very little capacitance data for liquid oxides [13,15,16]. This makes direct application of the solid-liquid interface models difficult. Also, if these data are to be used to confirm which model is most appropriate for use in liquid oxide-refractory systems then an accurate description of the contact area between the liquid oxide and refractory is required. This contact area is more easily described for dense refractories. The confirmation that an applied voltage across a refractory-liquid oxide interface for dense MgO affects the MgO corrosion rate is an essential first step in developing and understanding of this important phenomenon.

## 4. Conclusions

The effect of an applied voltage across the solid-liquid oxide interface on the dissolution rate of dense MgO refractory has been studied for two liquid oxide compositions in the CaO-SiO<sub>2</sub>-Al<sub>2</sub>O<sub>3</sub> and CaO-SiO<sub>2</sub>-Fe<sub>2</sub>O<sub>3</sub>-FeO-MgO liquid oxide systems respectively. It was found that the MgO corrosion in the CaO-SiO<sub>2</sub>-Al<sub>2</sub>O<sub>3</sub> based liquid oxide proved insensitive to an applied voltage over the voltage range -0.5 to 0.3 V. The corrosion rate of the dense MgO in the CaO-SiO<sub>2</sub>-Fe<sub>2</sub>O<sub>3</sub>-FeO-MgO proved to be effected by an applied voltage, showing a maximum corrosion rate at -0.45 V. This effect has been explained by considering the consequences of an applied voltage on the rate of Marangoni flow at the liquid oxide-refractory-gas interface and in turn, the flow effect on the rate of the mass transfer controlled MgO dissolution reaction.

The fact that the dissolution rate of a refractory can be altered by the application of a voltage across the refractory-liquid oxide interface offers a potential new tool to control refractory losses in liquid metal processing.

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